

## Chapter-2

### Worksheet-1

Q.1 What is the name of the metal which exists in liquid state at room temperature?

- (a) Sodium
- (b) Potassium
- (c) Mercury
- (d) Bromine

Q.2 When the liquid is spun rapidly, the denser particles are forced to the bottom and the lighter particles stay at the top. This principle is used in:

- (a) Centrifugation
- (b) Fractional distillation
- (c) Evaporation
- (d) Tunneling

Q.3 What is the name of the metal which exists in liquid state at room temperature?

- (a) Mercury
- (b) Bromine
- (c) Sodium
- (d) Potassium

Q.4 Which of the following elements is not a metalloid?

- (a) Boron
- (b) Silicon
- (c) Germanium
- (d) Tungsten

Q.5 If we put camphor in an open container, its amount keeps on decreasing due to the phenomenon of

- (a) Evaporation
- (b) Precipitation
- (c) Condensation
- (d) Sublimation

Q.6 Heterogeneous mixture in which the solute particles do not dissolve and remain suspended throughout the solvent and the solute particles can be seen with the naked eye is known as:

- (a) Colloidal solution
- (b) Super saturated solution
- (c) Sublimation
- (d) Suspensions

Q.7 In tincture of iodine, find the solute and solvent?

- (a) alcohol is the solute and iodine is the solvent
- (b) iodine is the solute and alcohol is the solvent
- (c) any component can be considered as solute or solvent
- (d) tincture of iodine is not a solution

Q.8 The continuous zig-zag movement of colloidal particles in a dispersion medium is called

- (a) Dispersion
- (b) Tyndall effect
- (c) Brownian movement
- (d) Oscillation

Q.9 A pure substance which is made up of only one kind of atom and cannot be broken into two or more simpler substances by physical or chemical means is referred to as

- (a) a compound
- (b) an element
- (c) a molecule
- (d) a mixture

Q.10 Which of the following non-metal is a good conductor of electricity?

(a) Aluminium

(b) Silicon

(c) Graphite

(d) Gold

Q.11 What is pure substance? Give two examples of pure substances.

Q.12 What are the two types of pure substances? Give one example of each type?

Q.13 Explain why, hydrogen and oxygen are considered elements whereas water is not considered an element.

Q.14 Is air a mixture or a compound? Give three reasons for your answer.

Q.15 Give two reasons for supposing that water is a compound and not a mixture.

Q.16 Define a mixture. Give two point of evidence to show that sugar solution is a mixture.

Q.17 State two reasons for supposing that brass is a mixture and not a compound.

Q.18 “The properties of the product are different from of the constituents.” State whether this statement best describes an element, a compound a mixture.

Q.19 Name one element, one compound and one mixture.

Q.20 what is the major difference between a solution and an ordinary mixture?